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Maximum total = [95
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against (temperature in
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order decomposition of
into and O. a.State how
the rate constant, varies
with temperature, . [1]
b.Determine the
activation energy, , for
this reaction. [3] The

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° C with hydrogen to
form nitrogen and
water. All reactants and
products are in the
gaseous phase. The gas-

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phase decomposition of dinitrogen monoxide is considered to occur in two steps. The experimental rate expression for this reaction is rate . The conversion of into is an exothermic reaction which can be represented as follows.

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